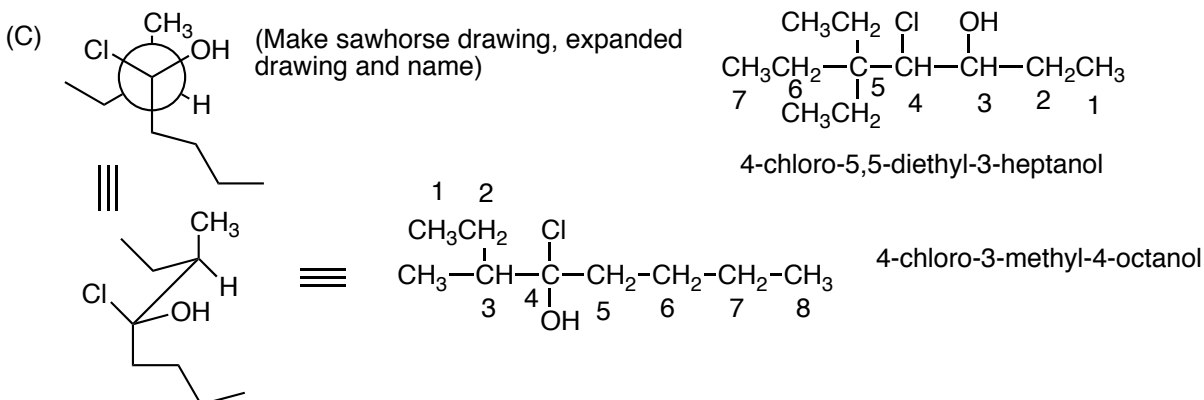
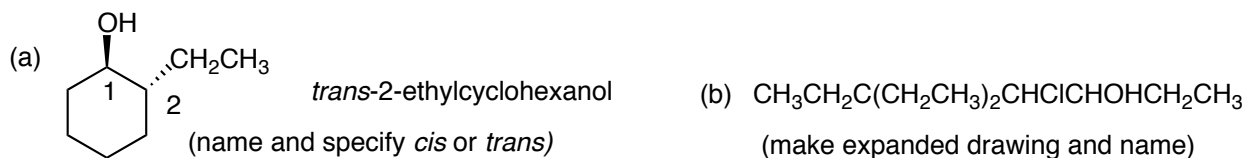
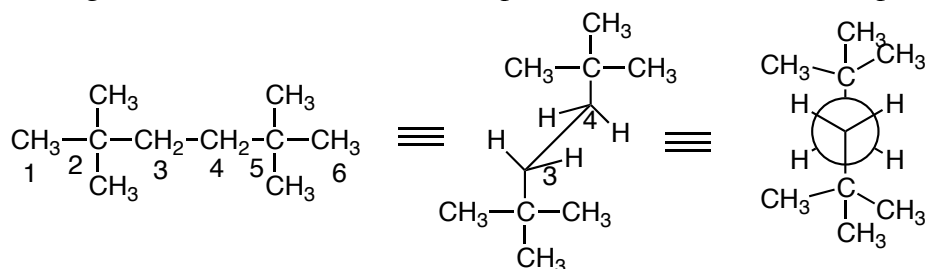
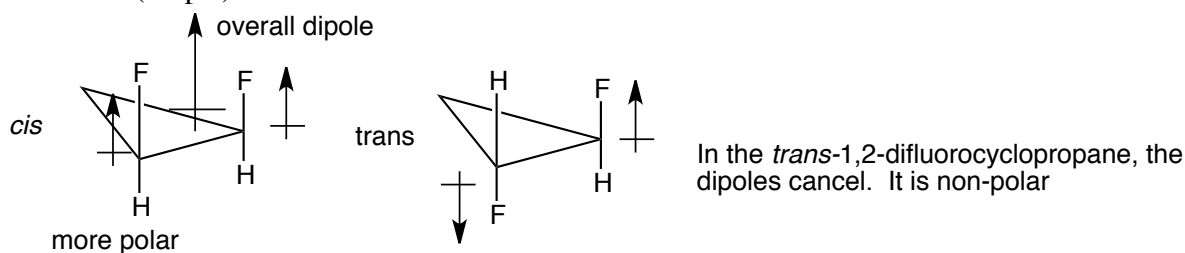
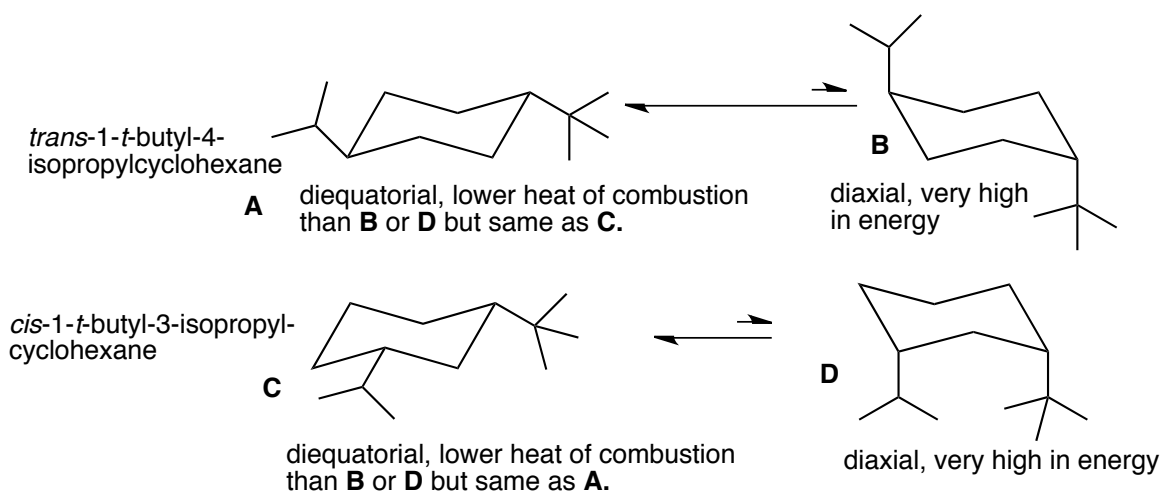


1. Name the following molecules. (15 pts)



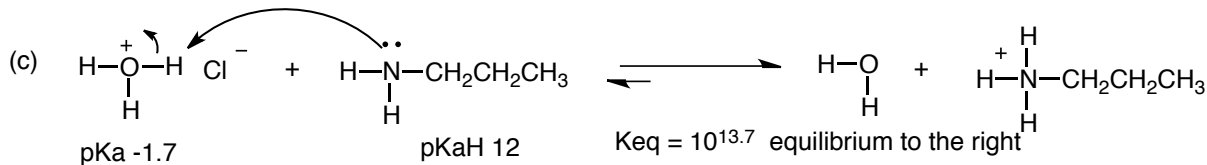
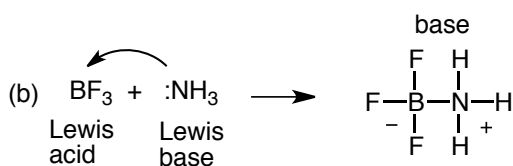
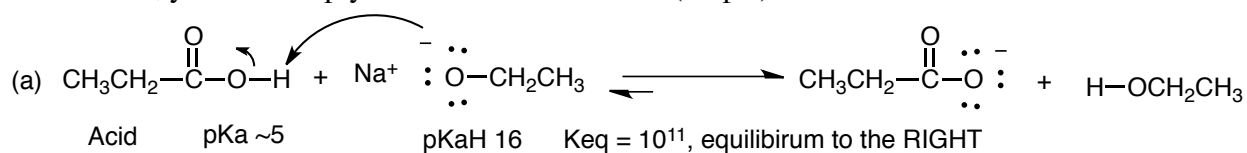
2. Draw a Newman projection of the low energy conformation of 2,2,5,5-tetramethylhexane looking down the C3-C4 bond and for partial credit first make an expanded drawing. (10 pts)

3. Which molecule is more polar, *cis*- or *trans*-1,2-difluorocyclopropane? Explain your answer by making careful three-dimensional drawings of each molecule, showing the relationship between the substituents, the individual dipole moments and the overall dipole moment for each molecule. (10 pts)4. Which molecule would have the lower heat of combustion, *trans*-1-*t*-butyl-4-isopropylcyclohexane or *cis*-1-*t*-butyl-3-isopropylcyclohexane? To answer this question you must make careful three-dimensional chair drawings of both molecules, showing both conformations and identifying in each case which conformation would have the lower heat of combustion and then finally choosing the lowest overall heat of combustion. (15 pts)



So, overall, the *cis*- and *trans* are the same in energy.

5. For the following acid-base reactions, identify the acid and the base, give the product of the reaction, show the movement of the electrons using the arrow formalism, show all charges, and also calculate the equilibrium constant for the reaction. If the equilibrium constant is very unfavorable, you can simply state NO REACTION. (30 pts)



6. Draw 4 isomers of C_5H_{10} . (20 pts). **BONUS:** Draw 2 more.

